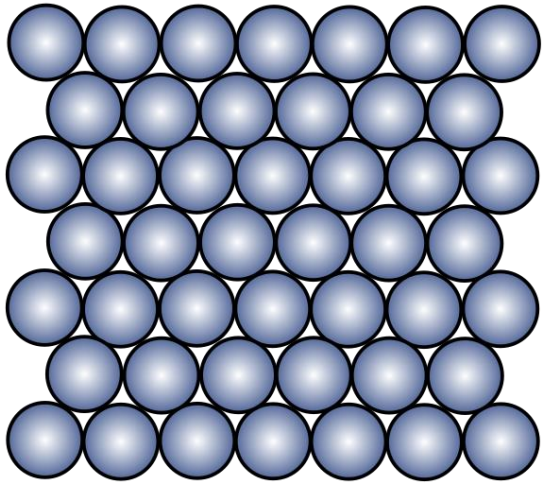
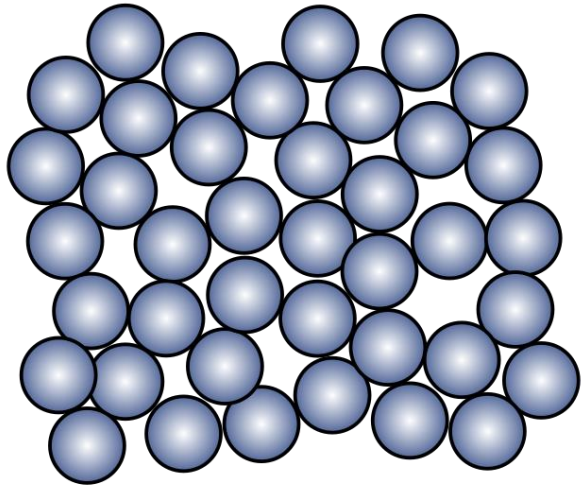
Structure of Matter

Dr Mohammad Abdur Rashid

Classification of solids

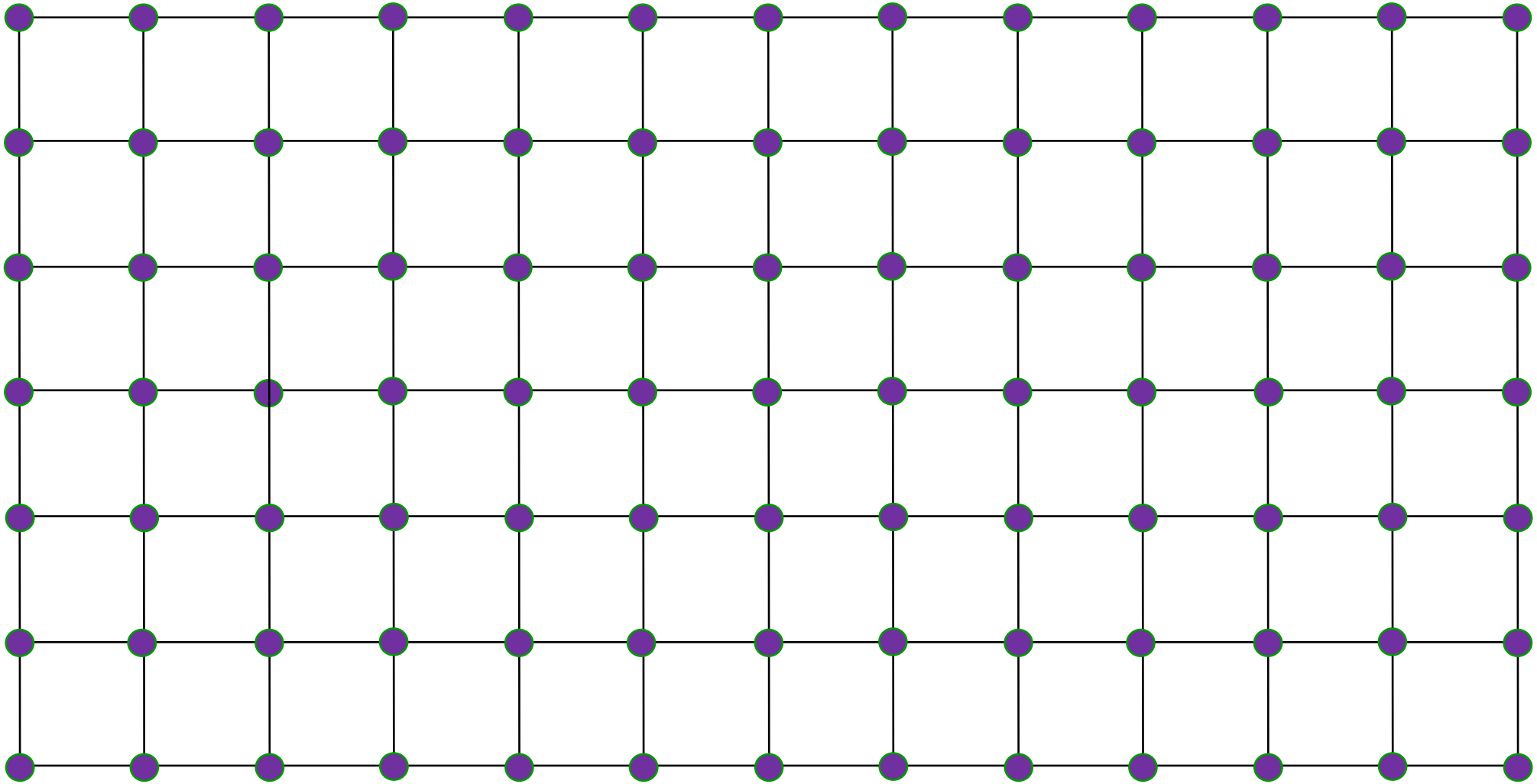
Solids are broadly classified into two types crystalline solids and amorphous solids.



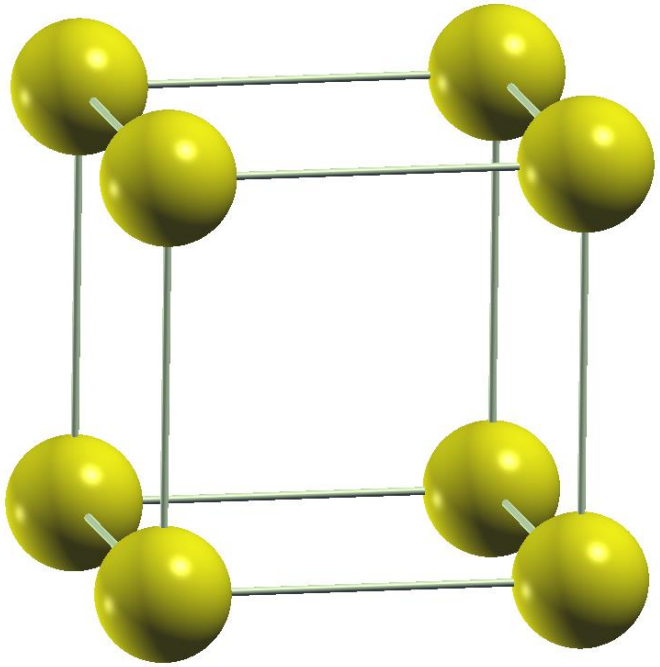


WIKIPEDIA
The Free Encyclopedia

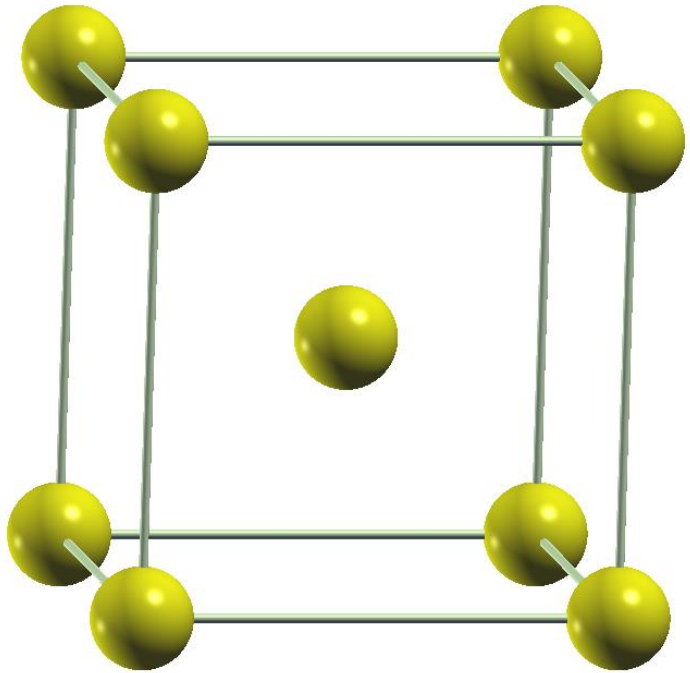
2D Crystal



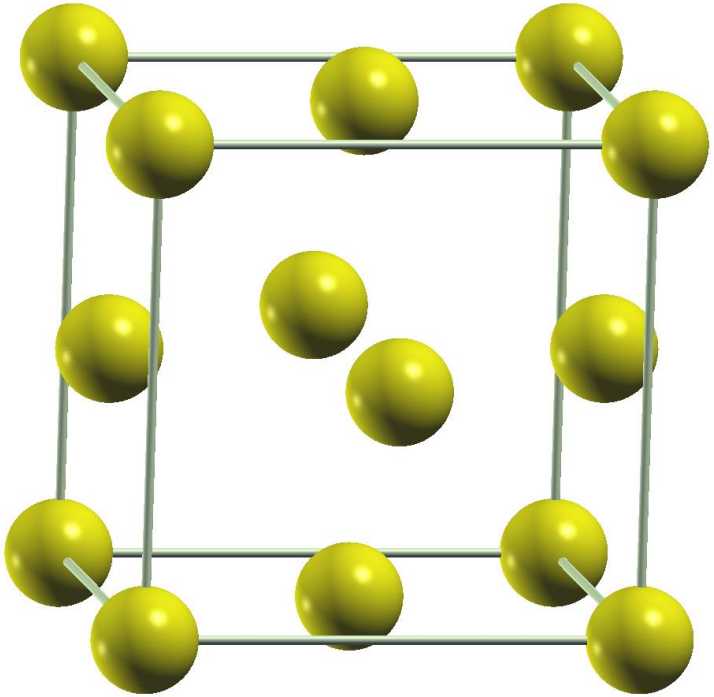
3D Crystal



Simple cubic

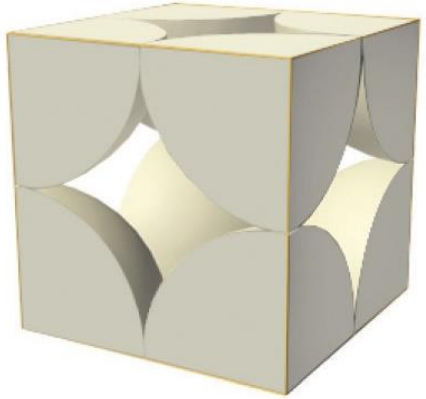


Body-centered cubic

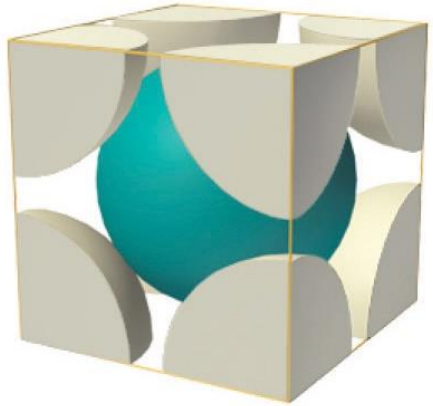


Face-centered cubic

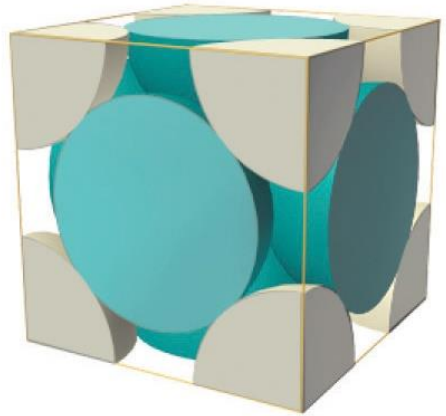
3D Crystal



Simple cubic



Body-centered cubic



Face-centered cubic

Unit cell

The Simplest Repeating Unit in a Crystal

The simplest repeating unit in a crystal is called a **unit cell**. Each unit cell is defined in terms of **lattice points,** the points in space about which the particles are free to vibrate in a crystal.

Bravais lattice

A Bravais lattice, studied by Auguste Bravais (1850), is an infinite array of discrete points in three dimensional space generated by a set of discrete translation operations described by:

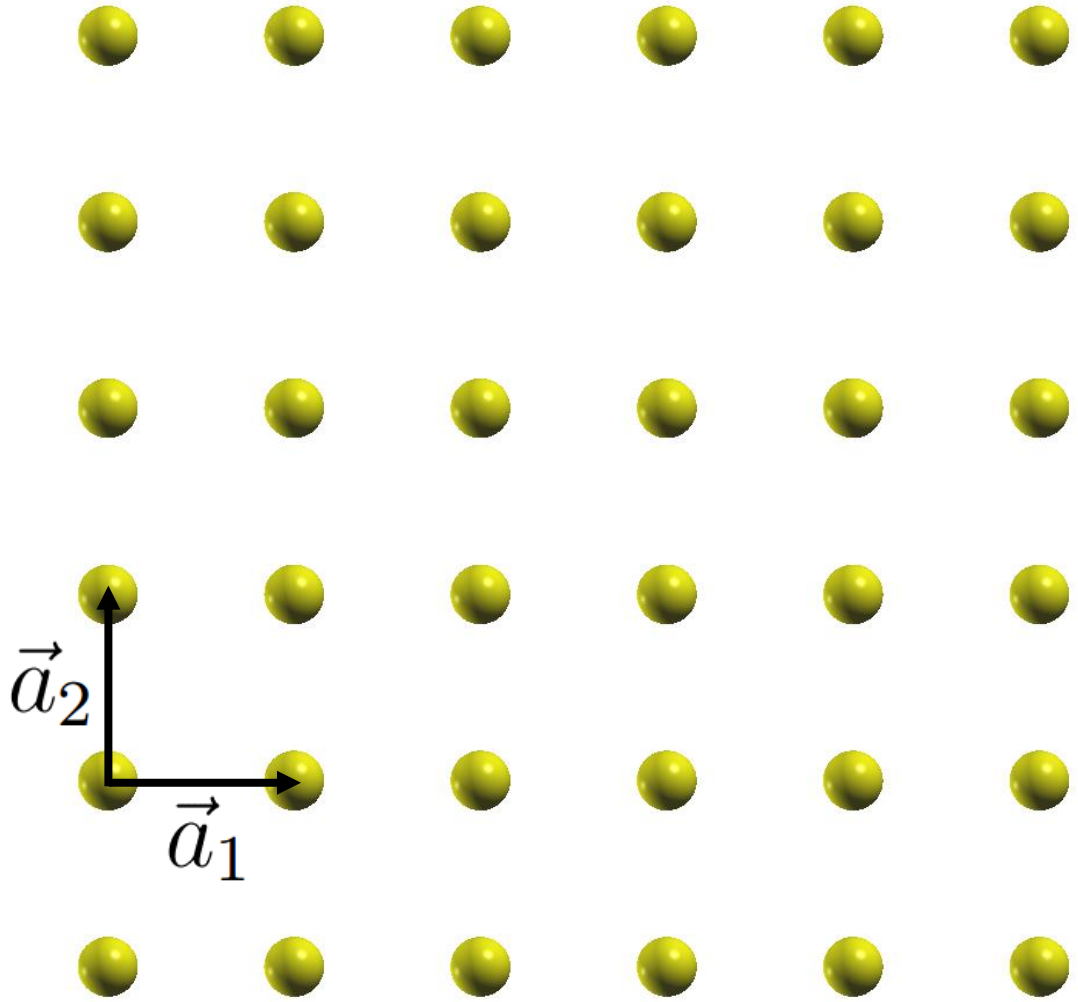
$$\vec{R} = n_1\vec{a}_1 + n_2\vec{a}_2 + n_3\vec{a}_3$$

Where n_i are any integers and $\vec{a}_i$ are known as the primitive vectors which lie in different directions and span the lattice. This discrete set of vectors must be closed under vector addition and subtraction. For any choice of position vector $\vec{R}$, the lattice looks exactly the same.

2D Crystal

$$\vec{R} = n_1\vec{a}_1 + n_2\vec{a}_2$$

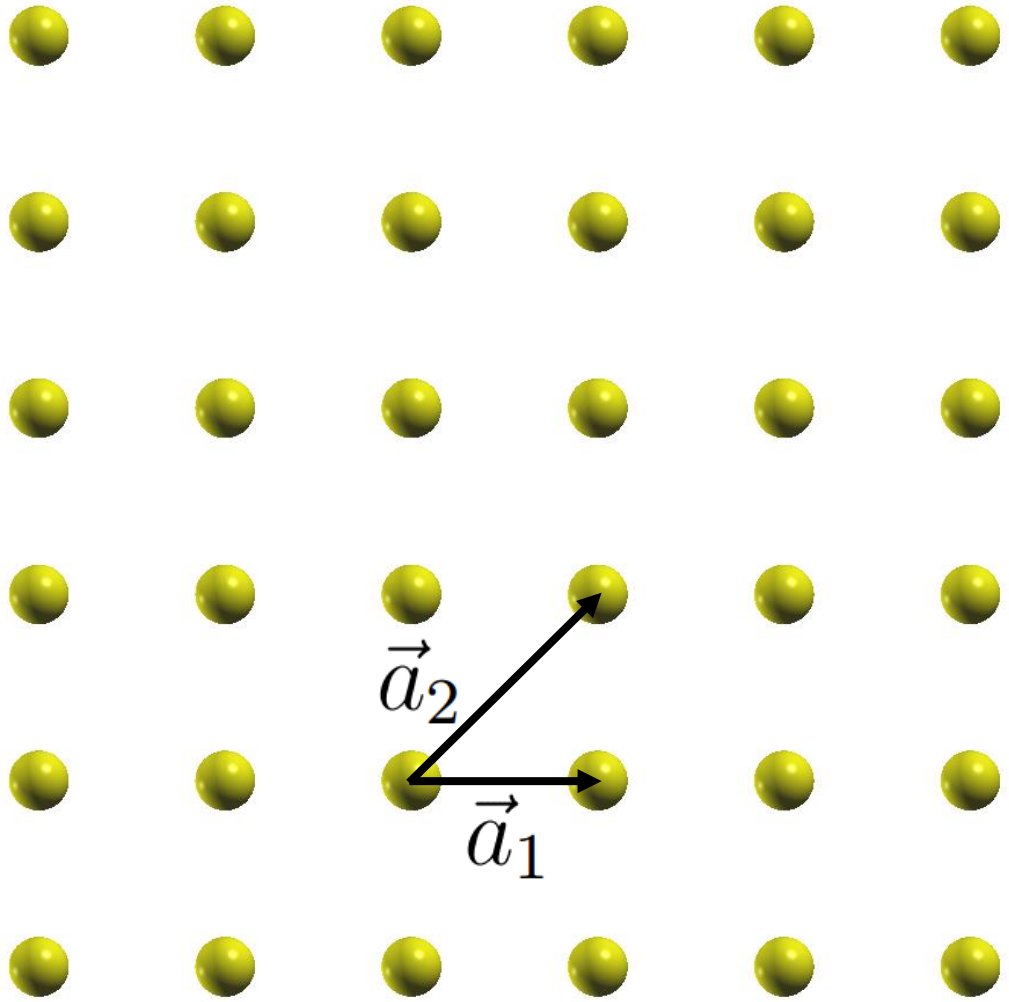
1	0
0	1
1	1
2	2



2D Crystal

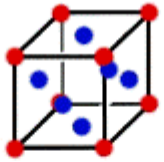
$$\vec{R} = n_1\vec{a}_1 + n_2\vec{a}_2$$

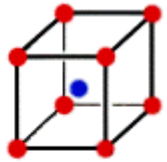
1	0
0	1
-1	1
1	1
-2	2

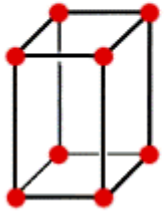


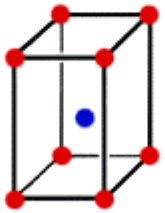
Bravais lattice in 3D

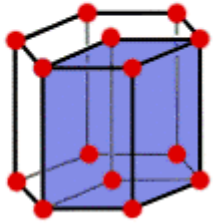

Simple cubic

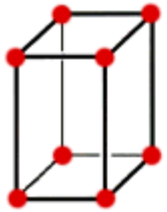

Face-centered cubic

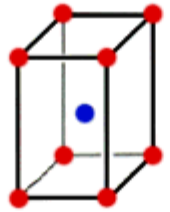

Body-centered cubic

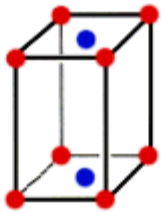

Simple tetragonal

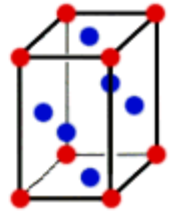

Body-centered tetragonal


Hexagonal

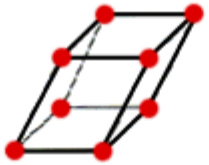

Simple orthorhombic

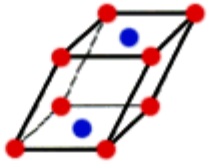

Body-centered orthorhombic

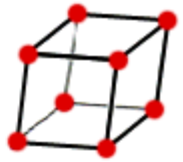

Base-centered orthorhombic


Face-centered orthorhombic


Rhombohedral


Simple Monoclinic


Base-centered monoclinic


Triclinic

epionelynx.wordpress.com

Bravais lattice

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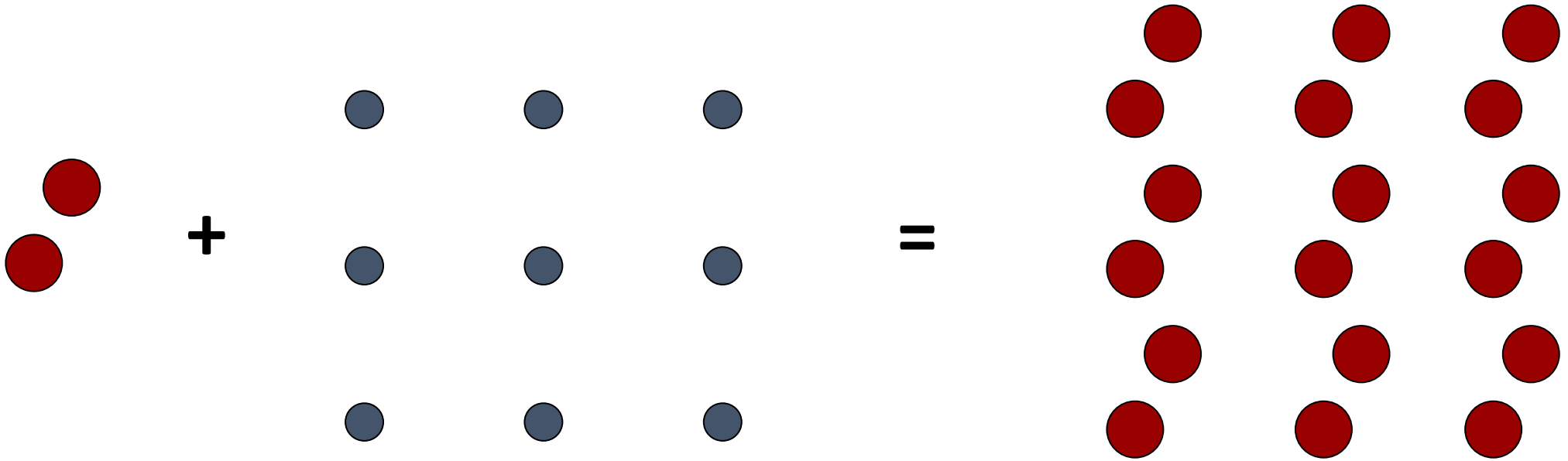
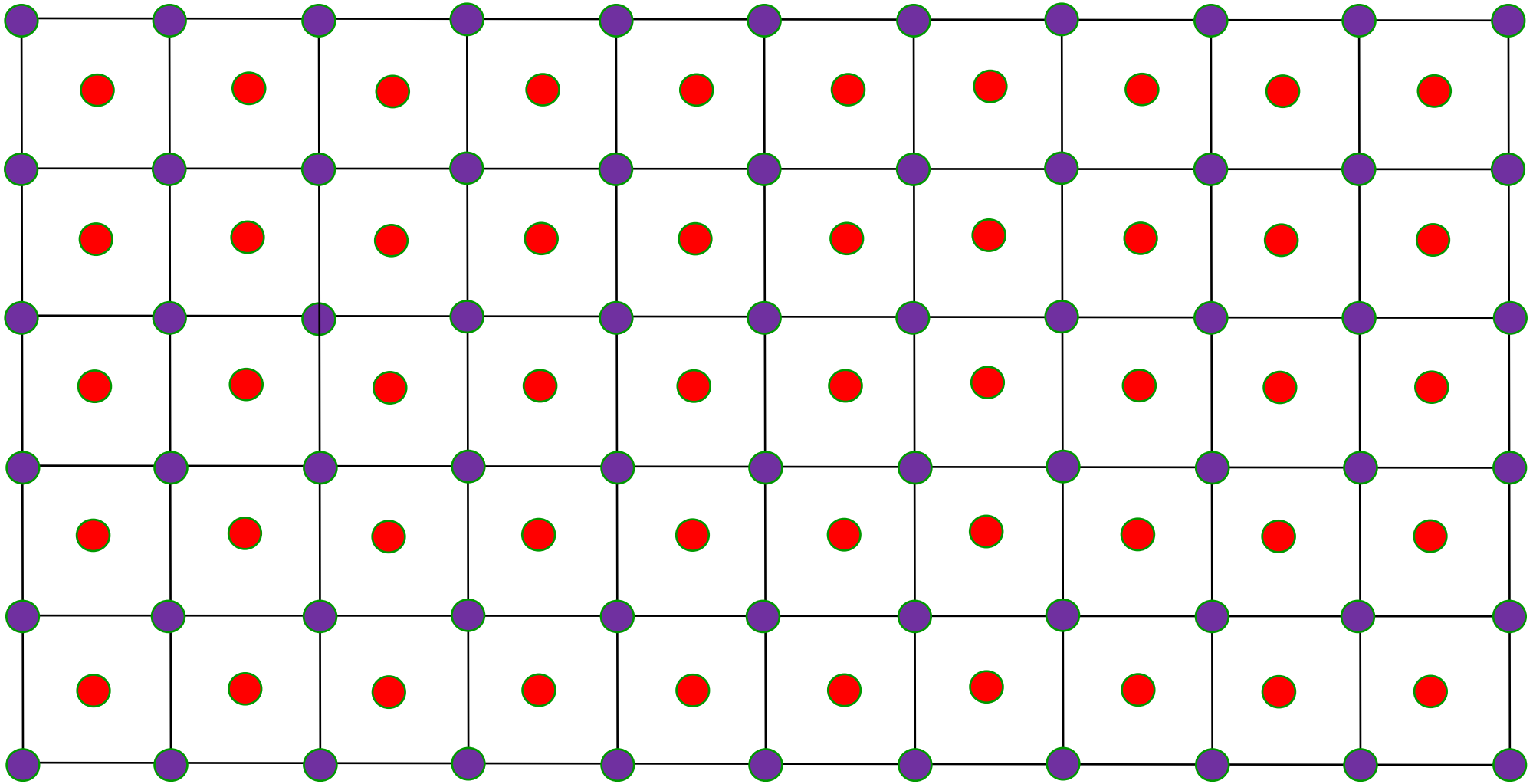
Where n_i are any integers and $\vec{a}_i$ are known as the primitive vectors which lie in different directions and span the lattice. This discrete set of vectors must be closed under vector addition and subtraction. For any choice of position vector $\vec{R}$, the lattice looks exactly the same.

Crystal structure

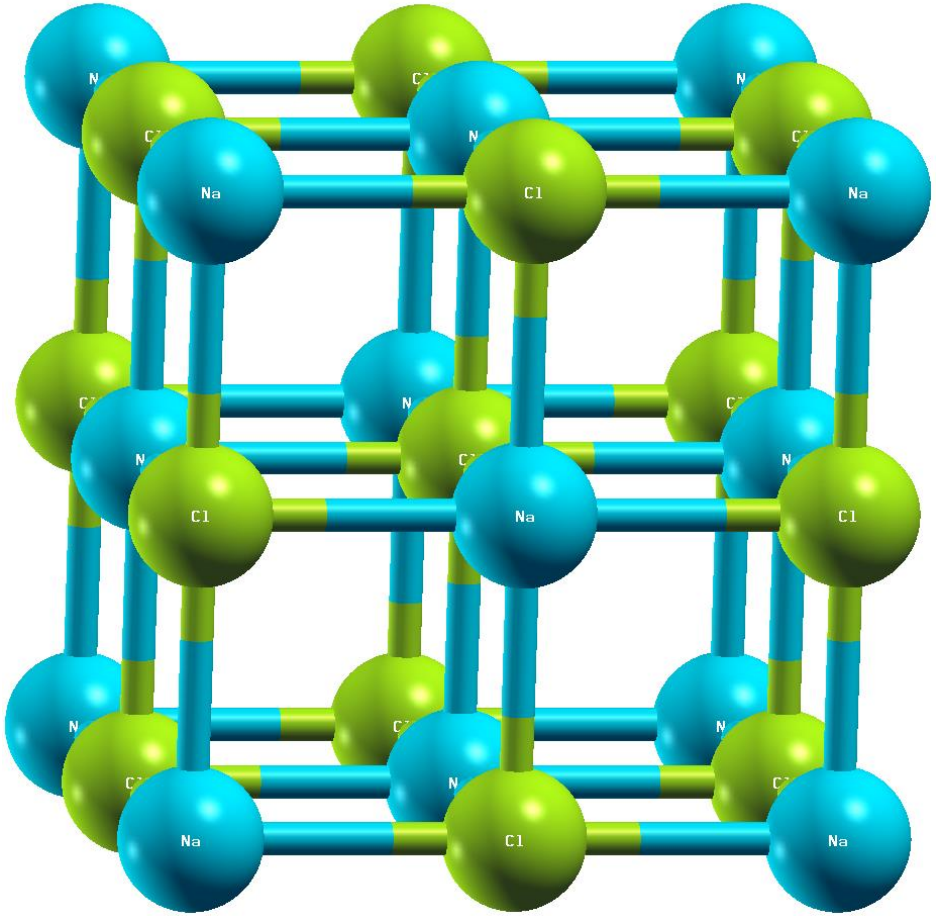
basis + lattice = crystal structure

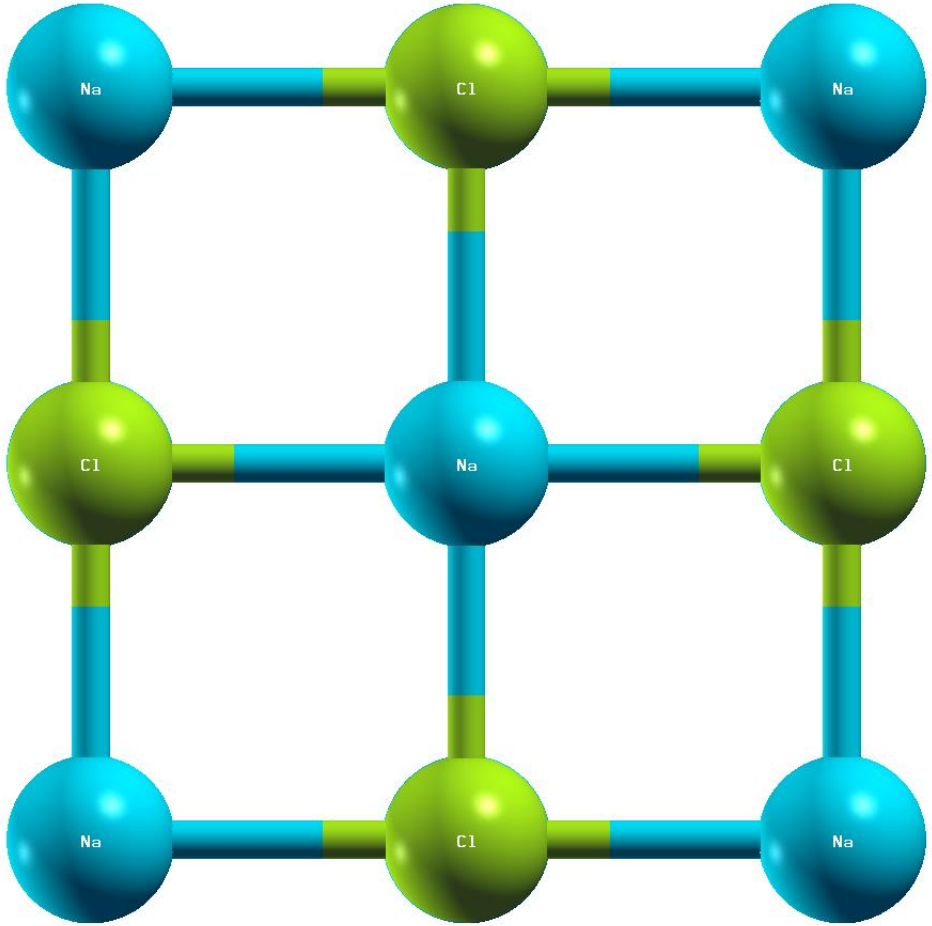


Crystal structure

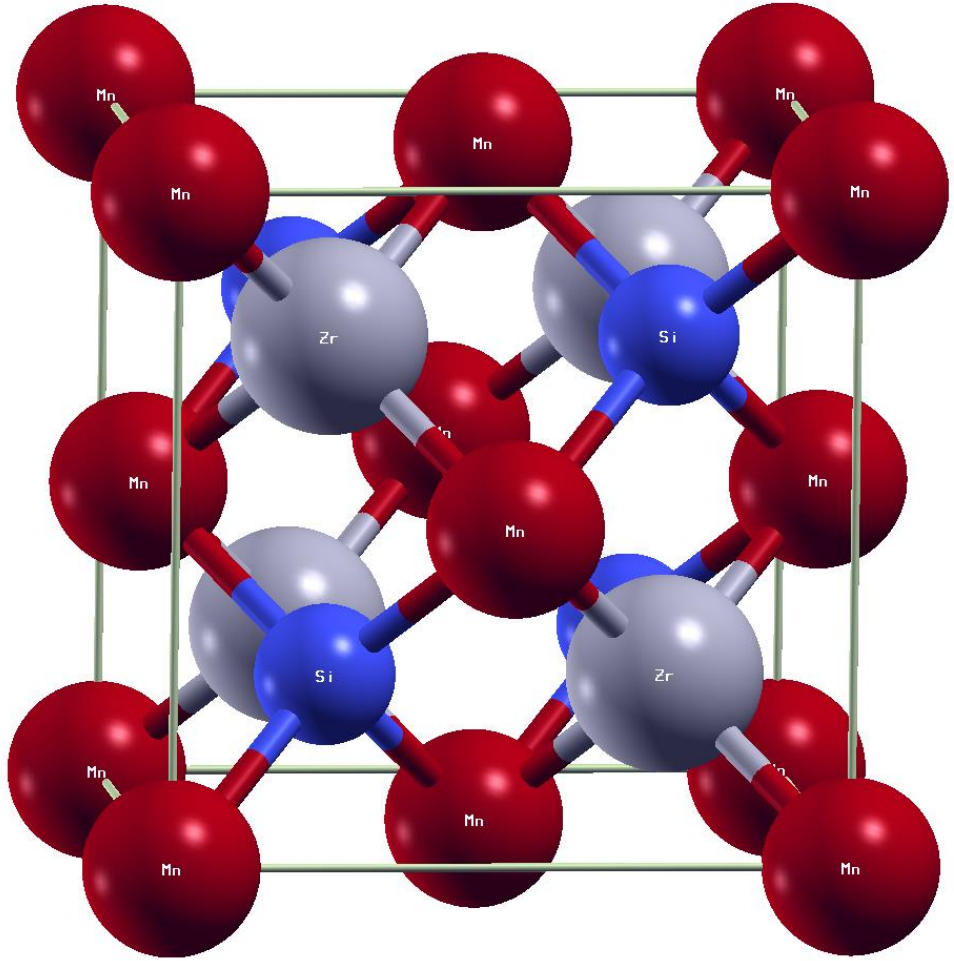


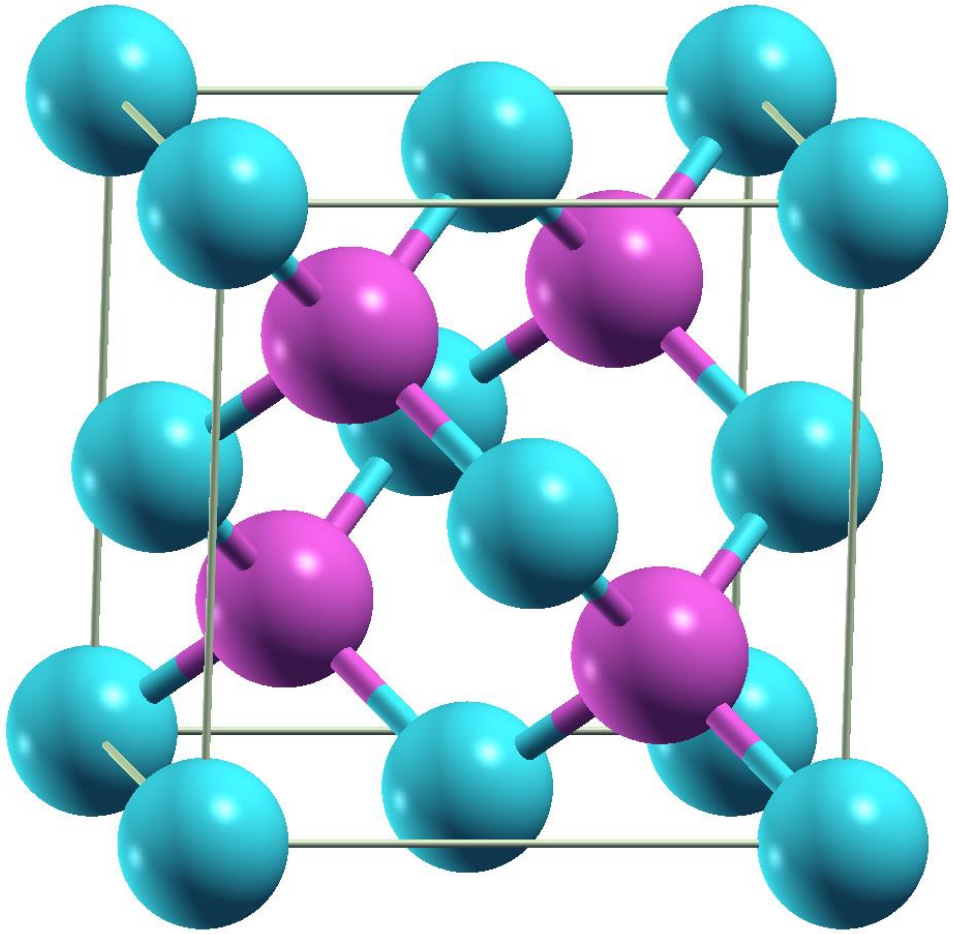
Crystal structure





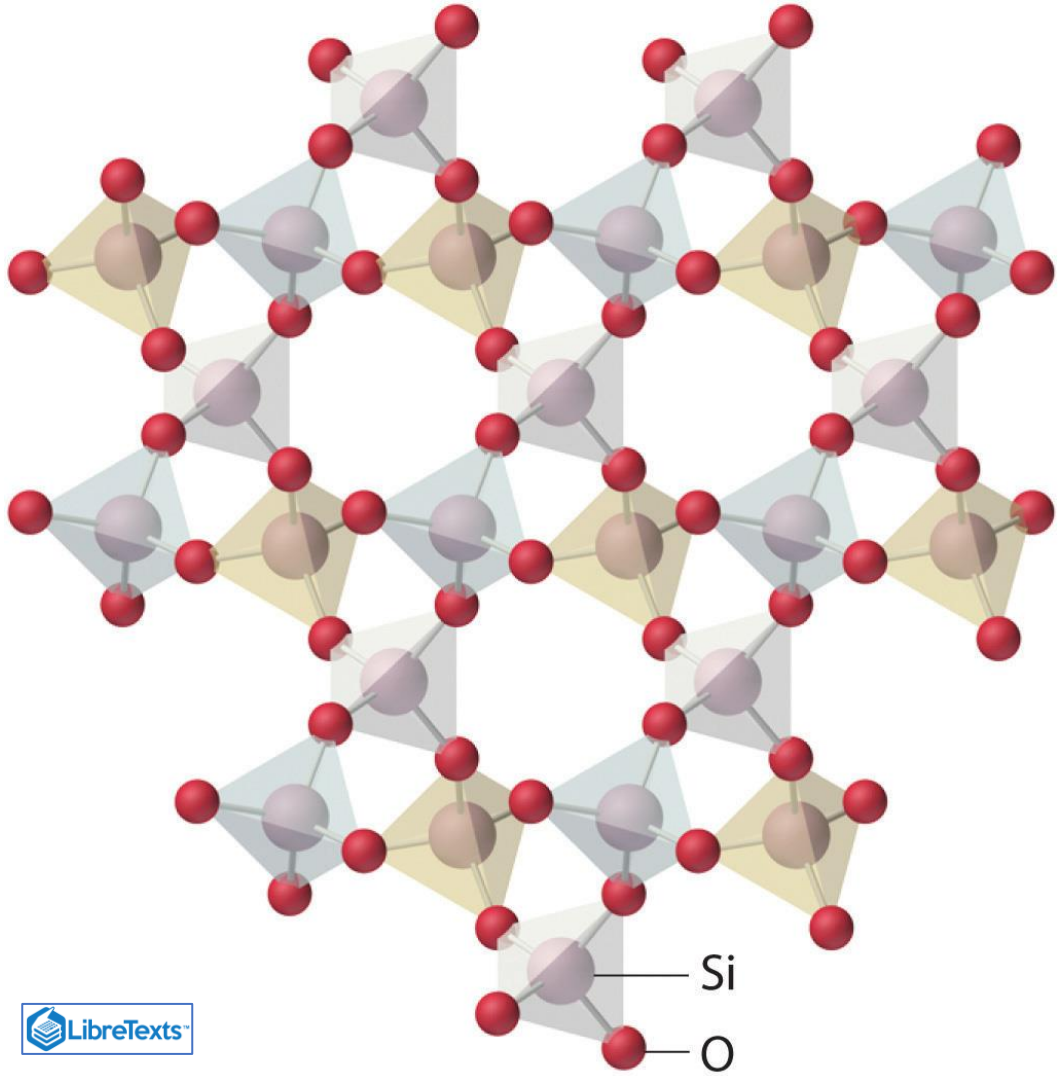
Crystal structure



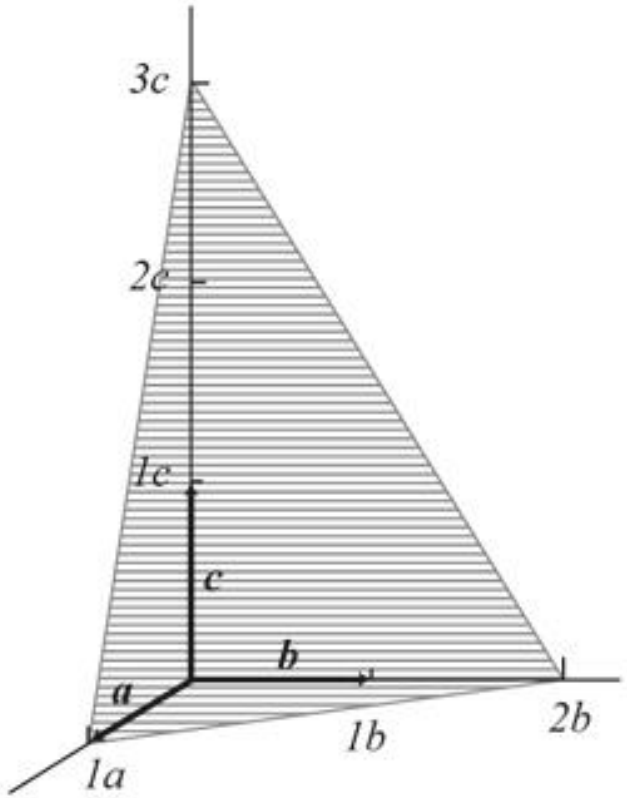


Crystal structure

The lattice of crystalline quartz (SiO_2)



Miller Indices



The Miller indices of a lattice plane are the coordinates of the shortest reciprocal lattice vector normal to the plane, with respect to a specific set of primitive reciprocal lattice vectors. Thus a plane with Miller indices ($h\ k\ l$), is normal to the reciprocal lattice vector

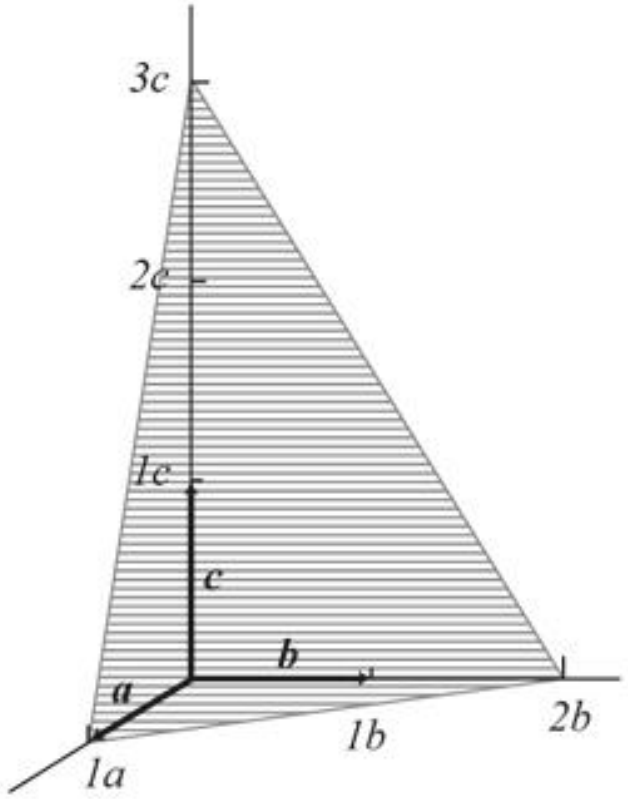
$$h\vec{b}_1 + k\vec{b}_2 + l\vec{b}_3$$

Miller indices are used to specify directions and planes.

The rules for Miller Indices:

- Determine the intercepts of the face along the crystallographic axes, *in terms of unit cell dimensions.*
- Take the reciprocals
- Clear fractions
- Reduce to lowest terms

$$(1, 2, 3) \rightarrow (\frac{1}{1} : \frac{1}{2} : \frac{1}{3}) \rightarrow (6 : 3 : 2) \rightarrow (632)$$



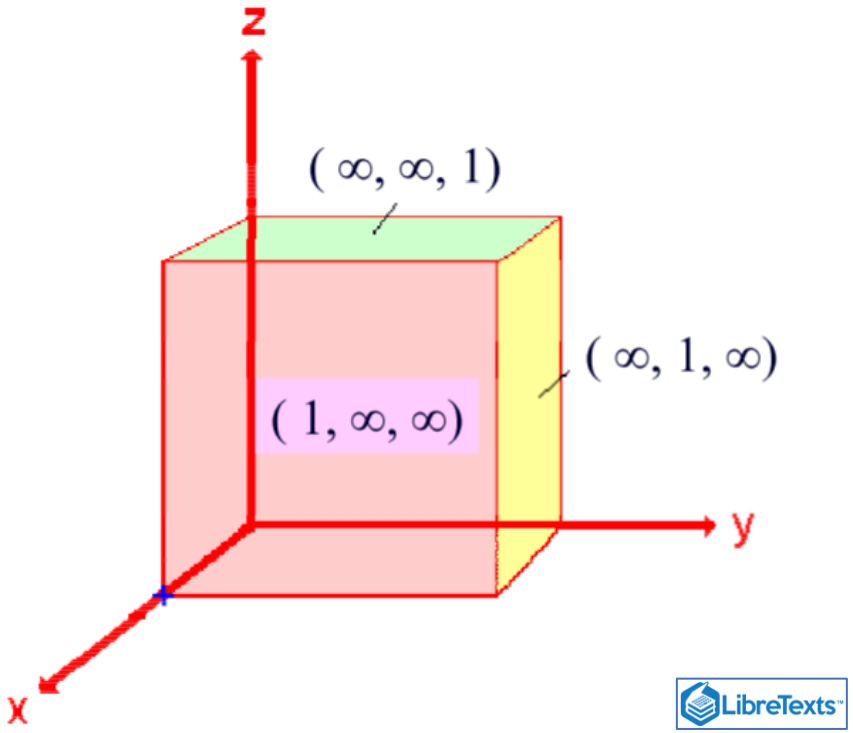
Miller Indices

Miller Indices are the reciprocals of the parameters of each crystal face.

For Pink Face: $\left(\frac{1}{1}, \frac{1}{\infty}, \frac{1}{\infty}\right) \rightarrow (100)$

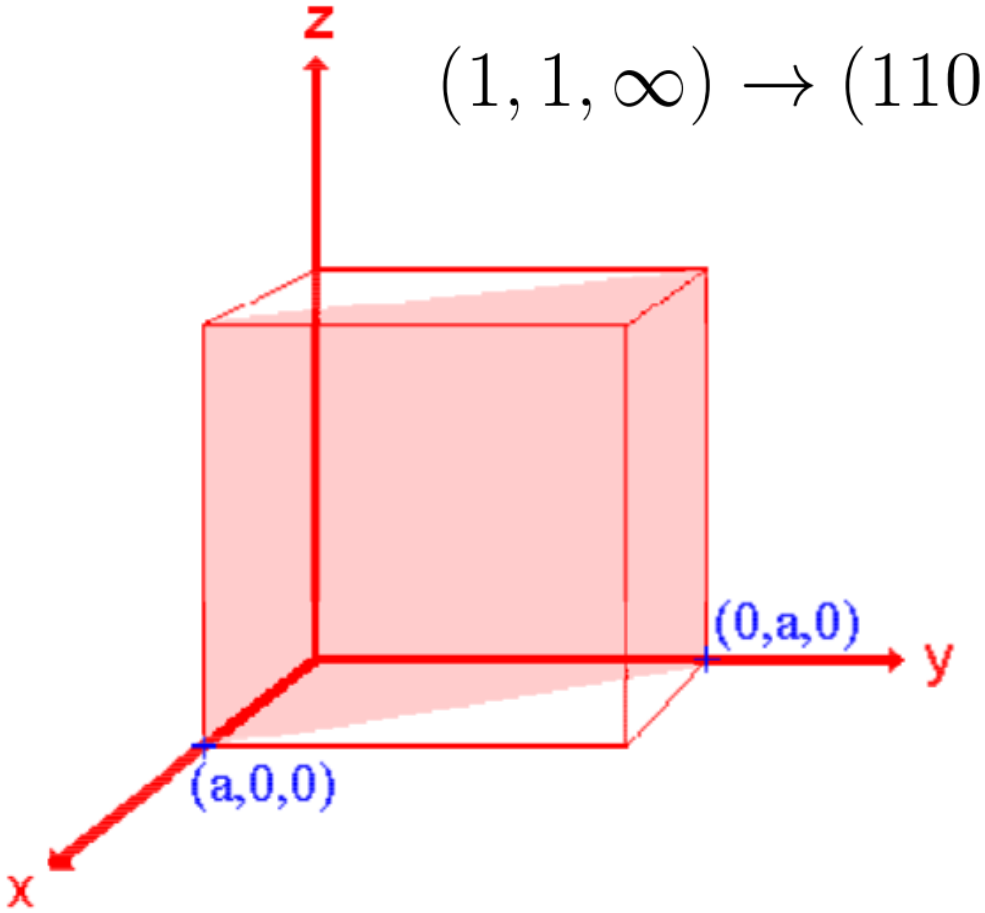
For Green Face: $\left(\frac{1}{\infty}, \frac{1}{\infty}, \frac{1}{1}\right) \rightarrow (001)$

For Yellow Face: $\left(\frac{1}{\infty}, \frac{1}{1}, \frac{1}{\infty}\right) \rightarrow (010)$

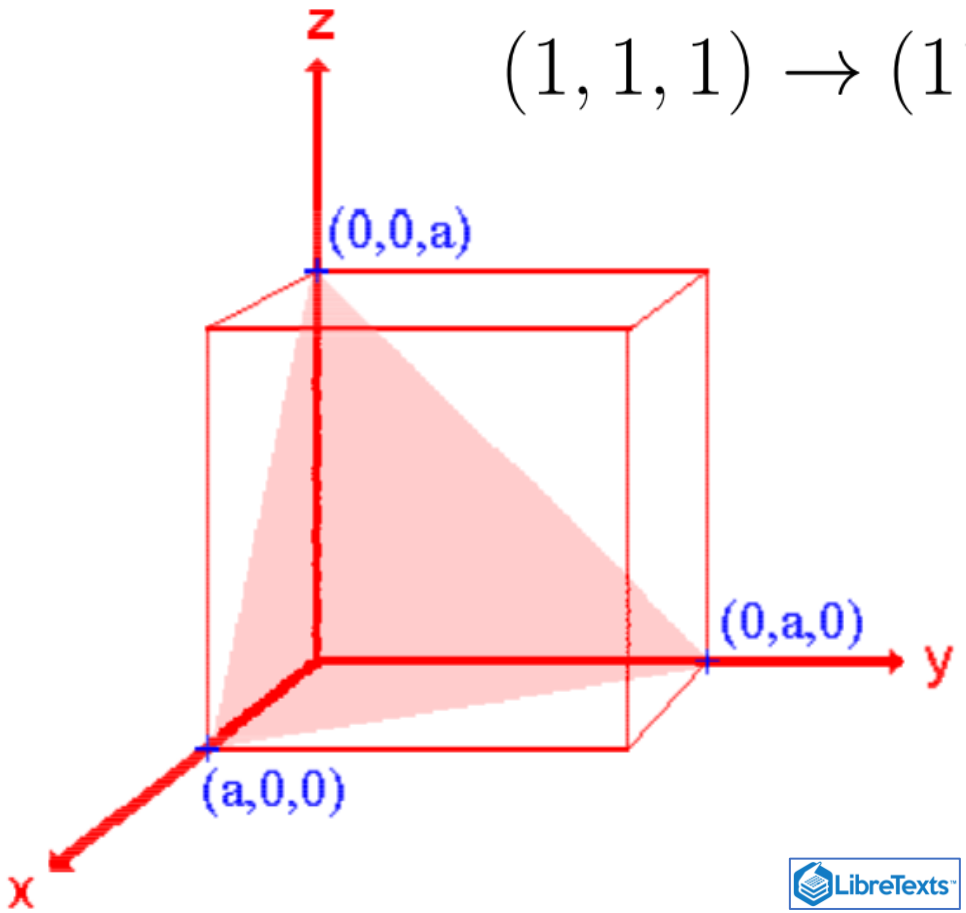


Miller Indices

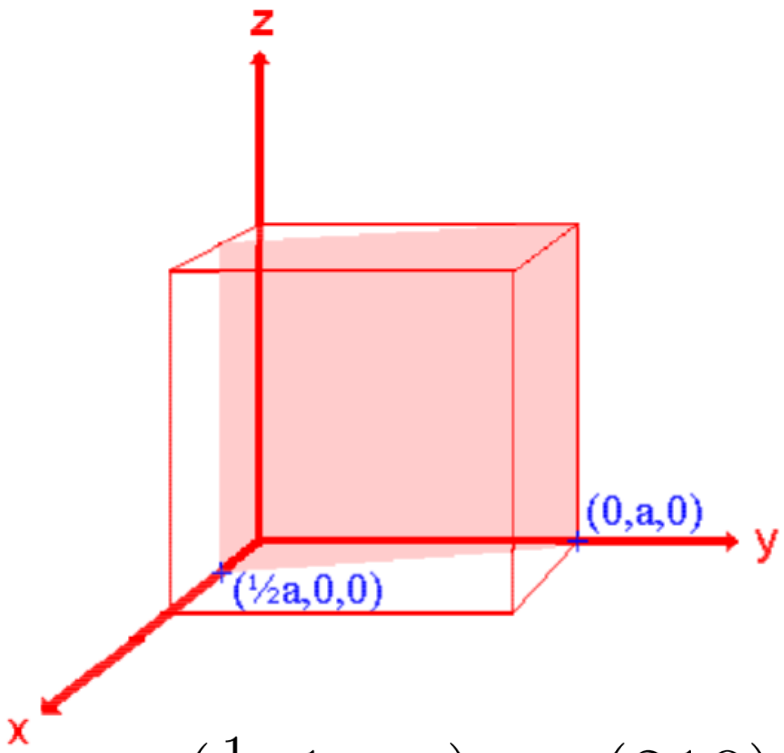
$(1, 1, \infty) \rightarrow (110)$



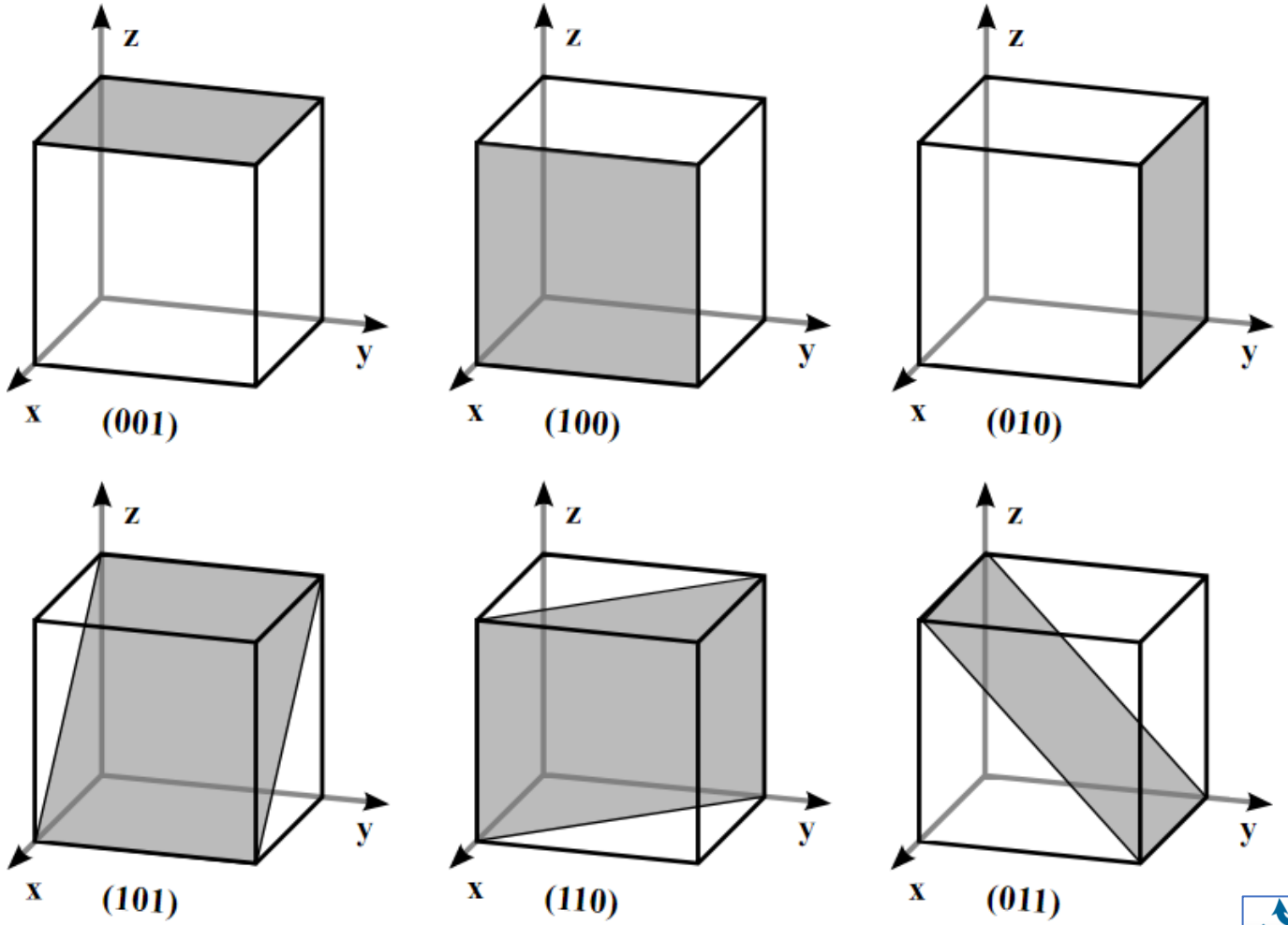
$(1, 1, 1) \rightarrow (111)$



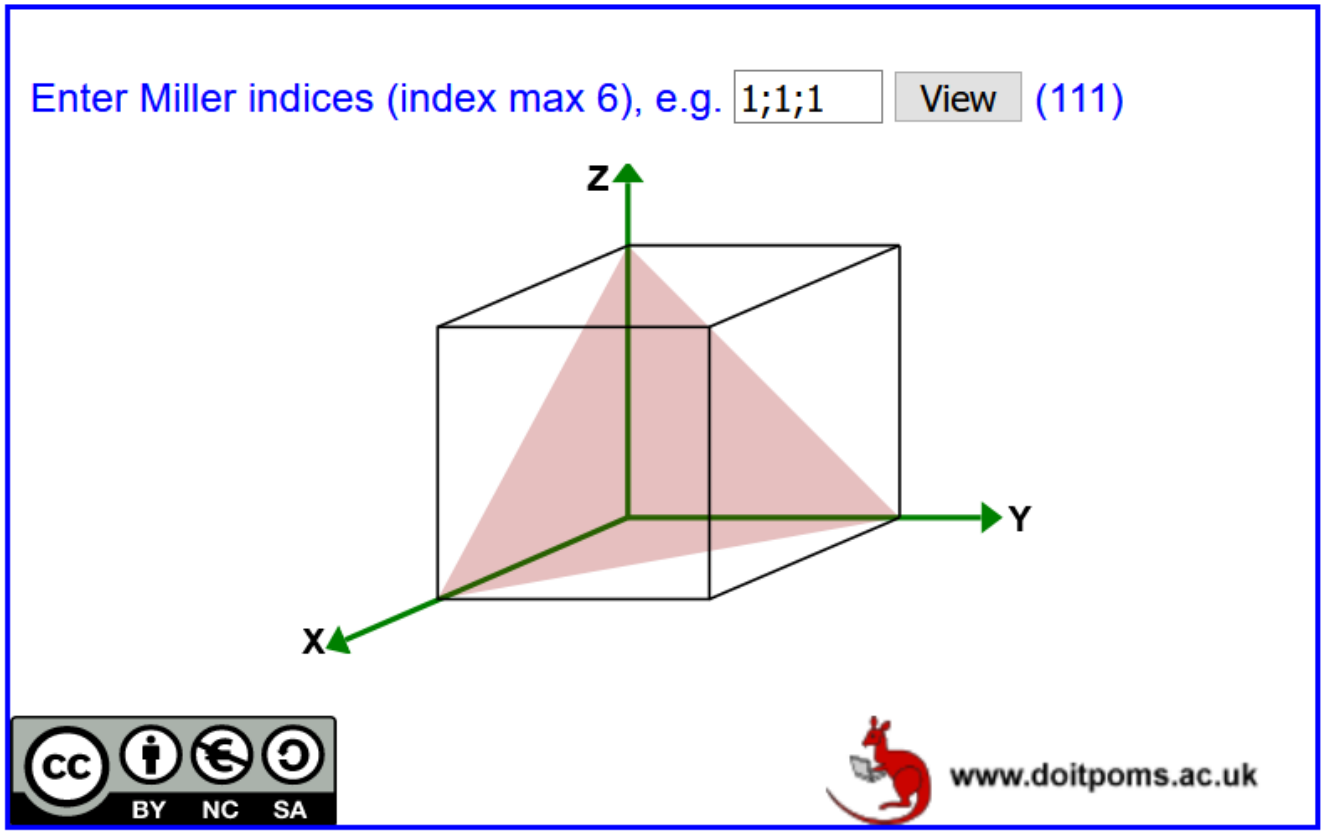
Miller Indices



$$\left(\frac{1}{2}, 1, \infty\right) \rightarrow (210)$$



Draw your own lattice planes

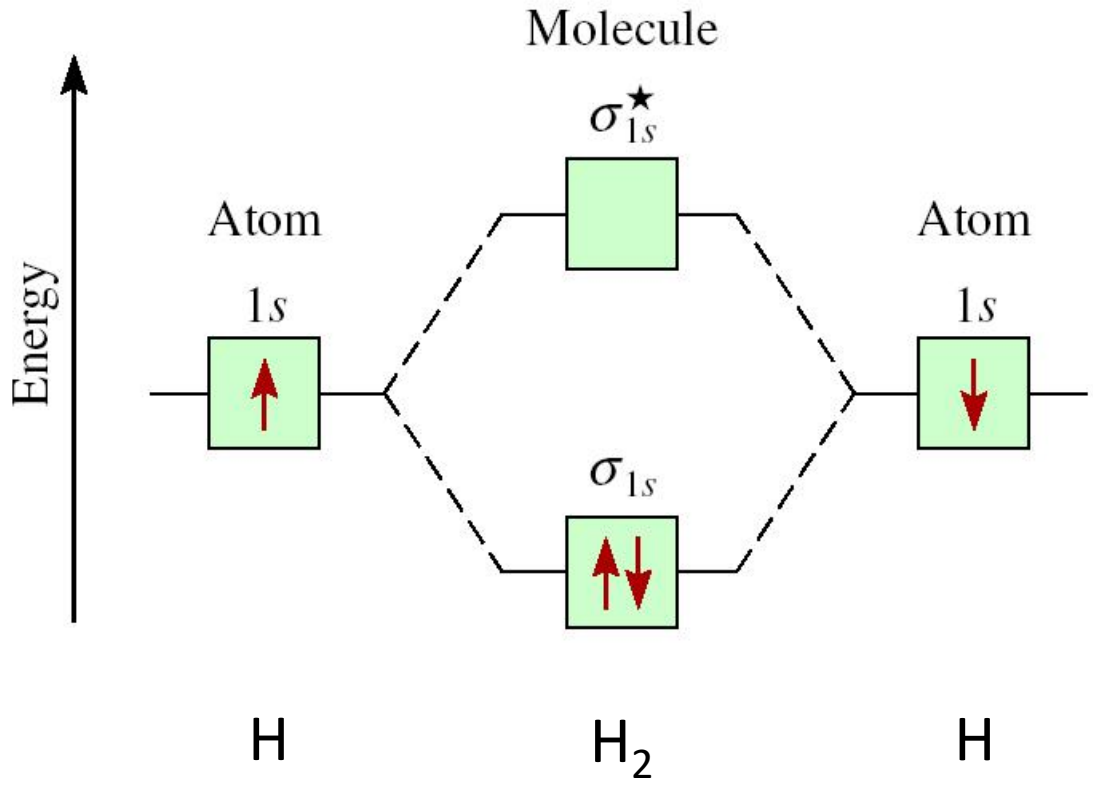


https://www.doitpoms.ac.uk/tlplib/miller_indices/lattice_draw.php

Formation of molecule

When two atoms are brought together

- Atomic energy levels split
- Molecular orbitals are formed



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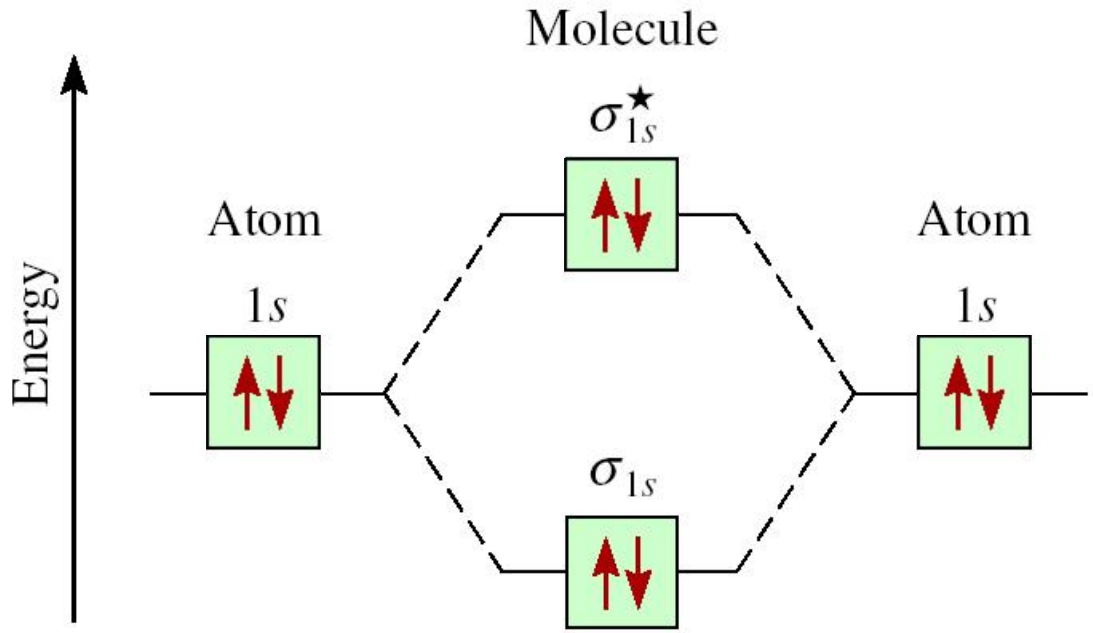

He-He system

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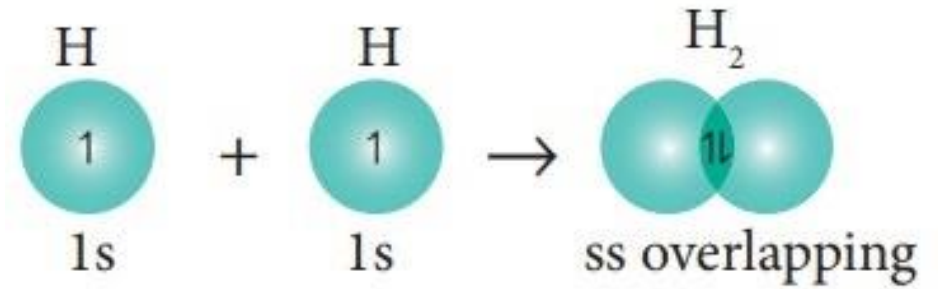
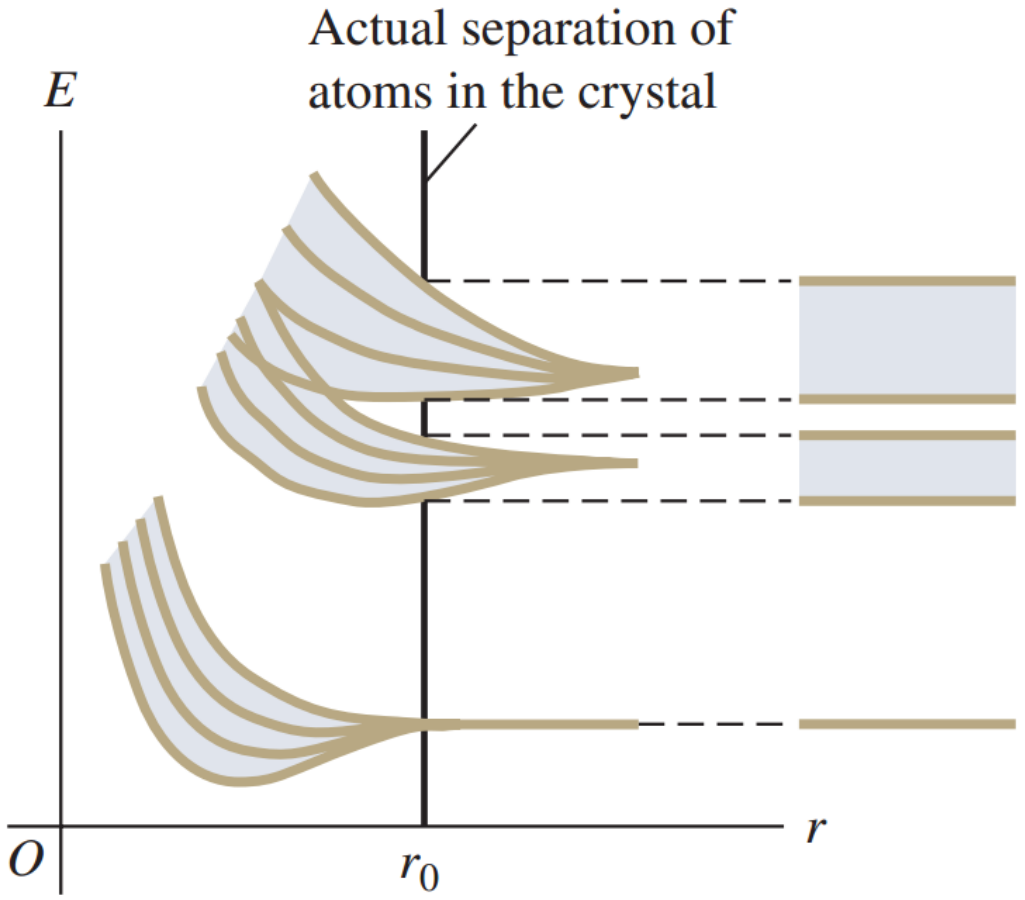
Band Theory of Solids



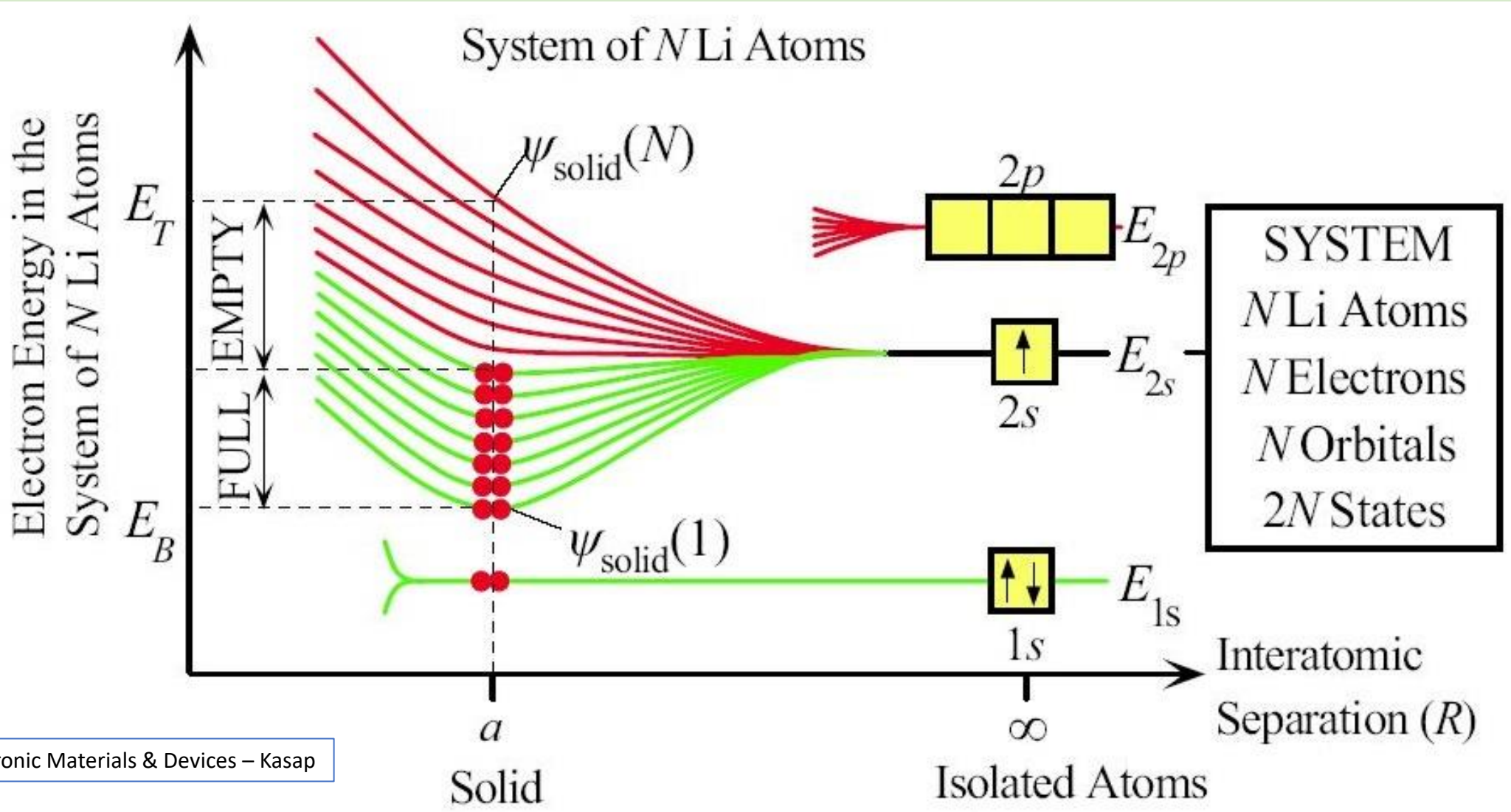
Formation of hydrogen molecule



brainkart.com

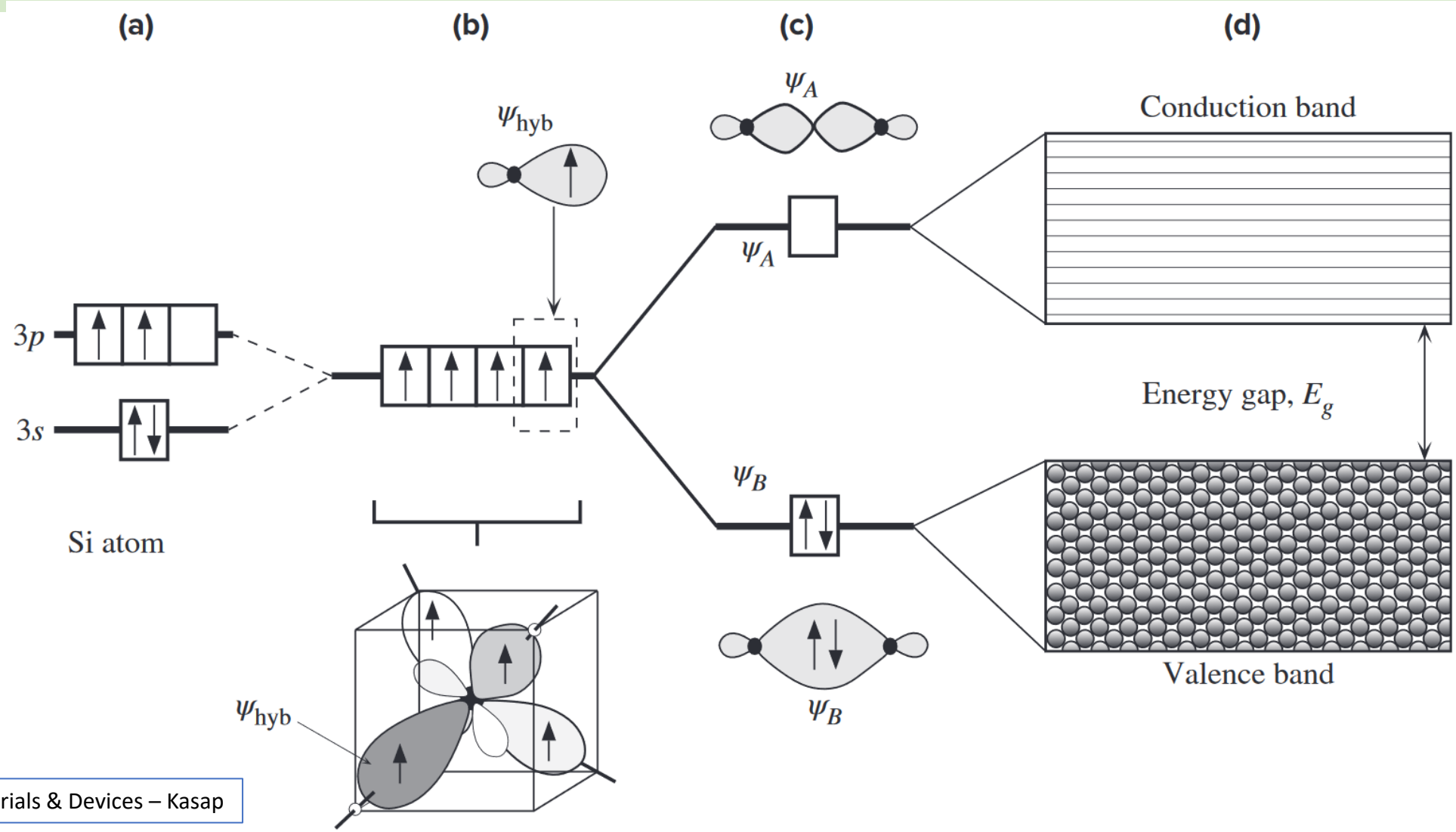
University Physics– Young, Freedman

Formation of solid - Lithium



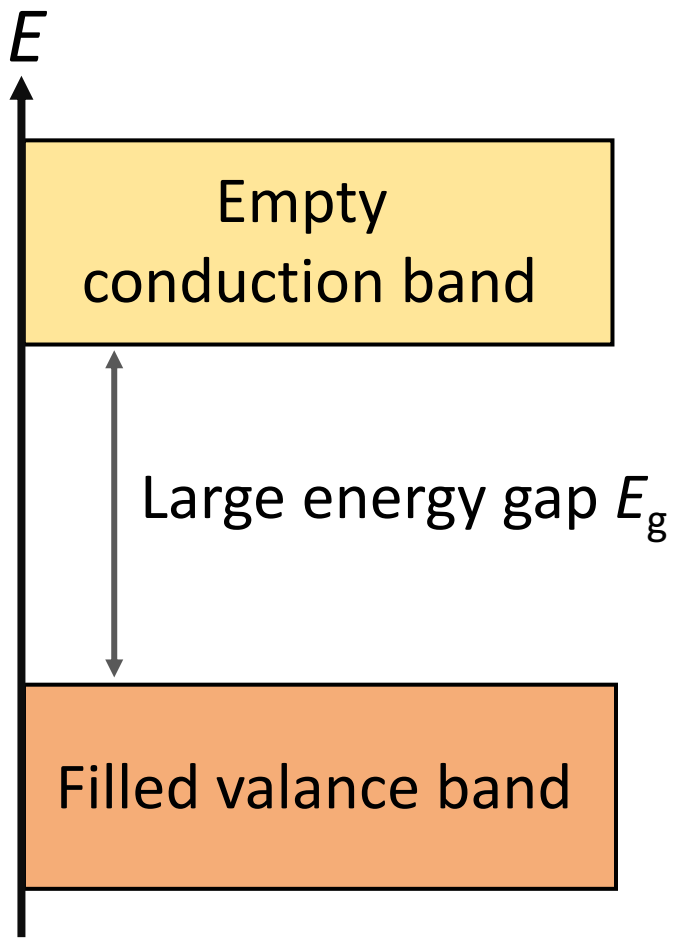
Electronic Materials & Devices – Kasap

Formation of solid – Silicon crystal

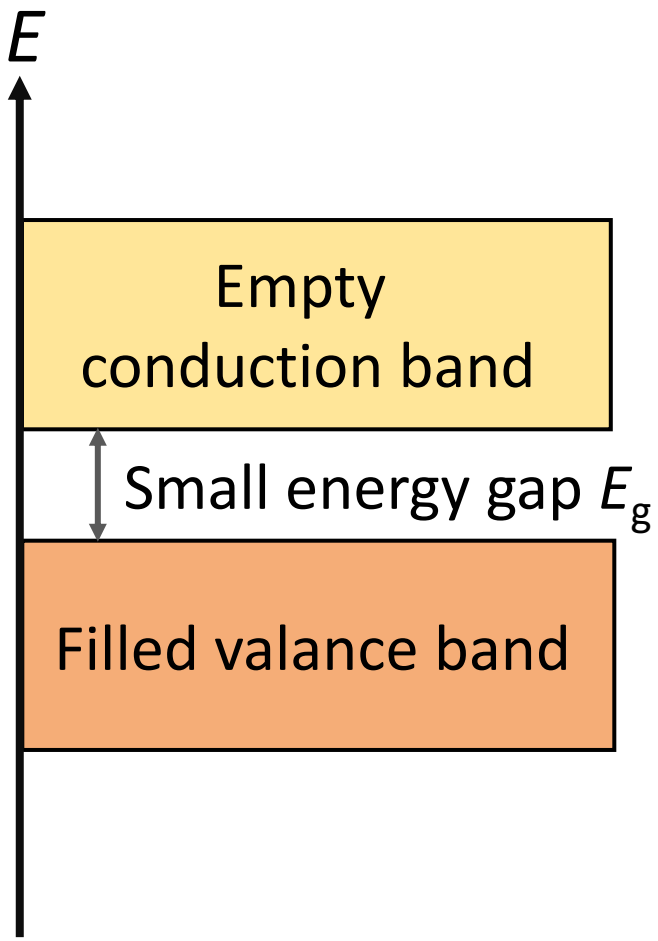


Electronic Materials & Devices – Kasap

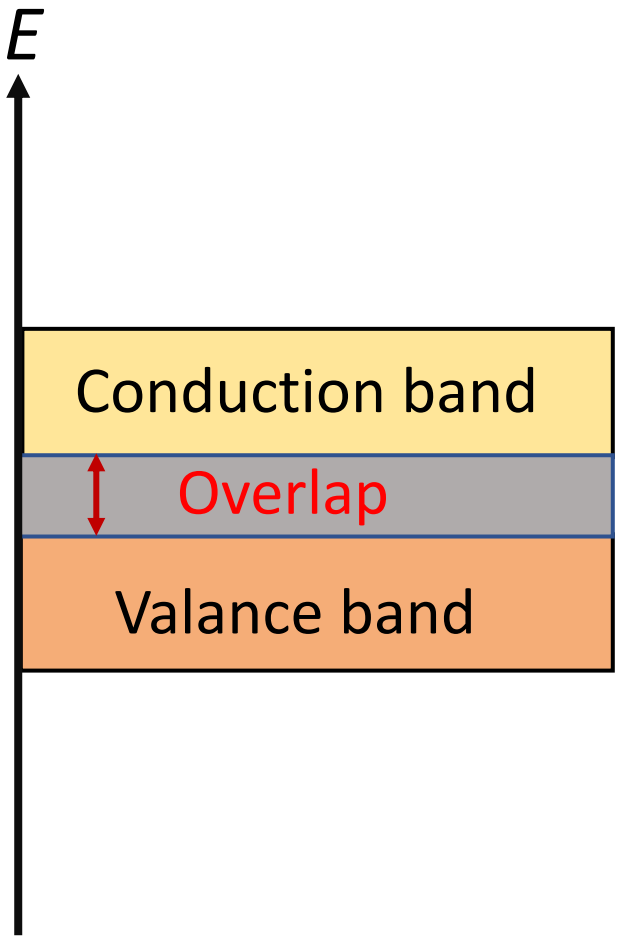
Energy Bands for Solids



Insulator

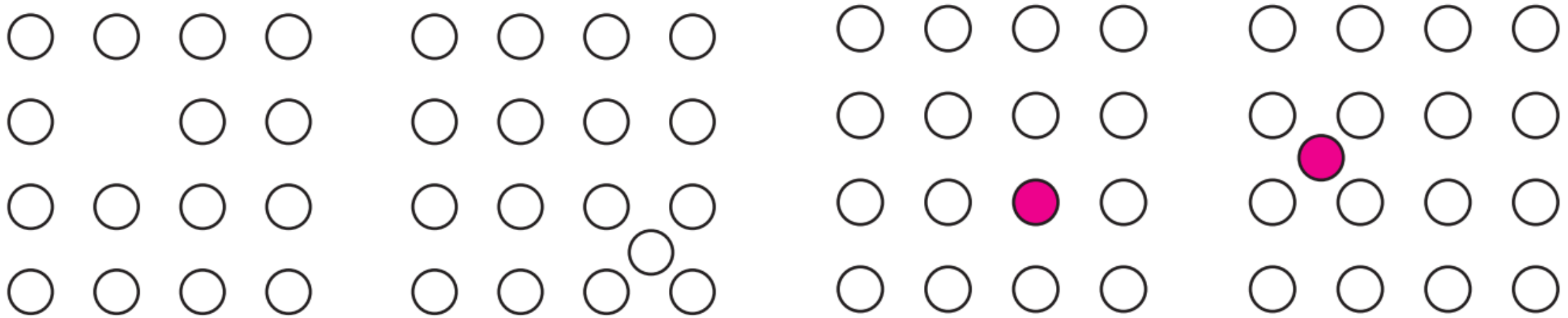


Semiconductor



Conductor

Crystal Defects



Point defects in a crystal. (*a*) Vacancy. (*b*) Interstitial. (*c*) Substitutional impurity. (*d*) Interstitial impurity.

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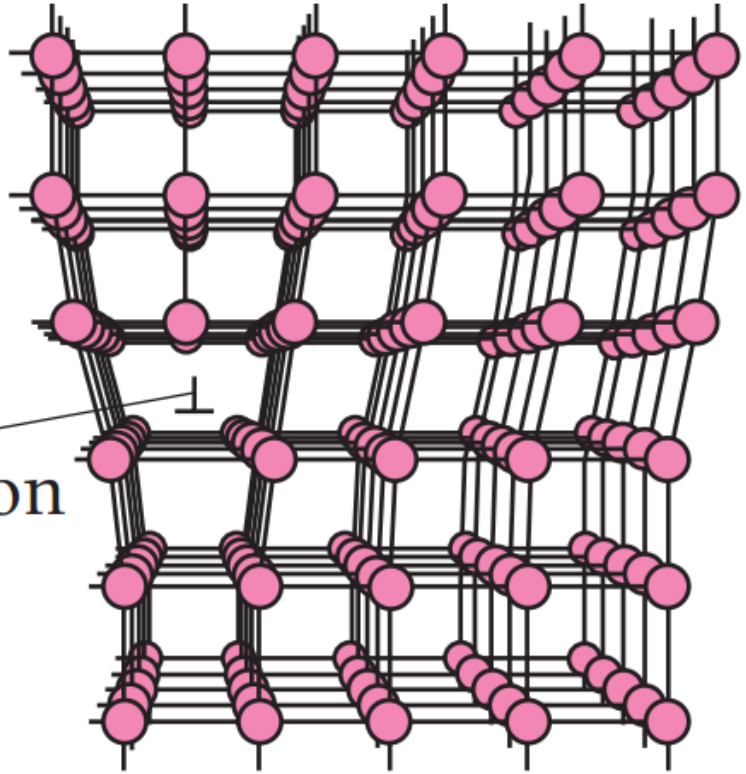
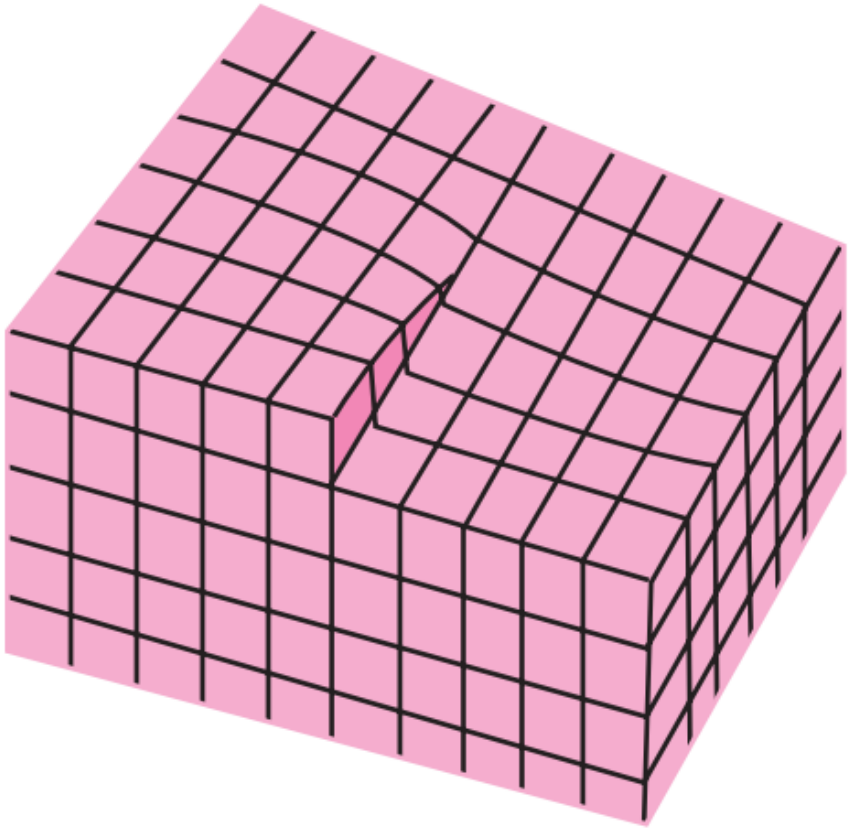
Crystal Defects





A screw dislocation.

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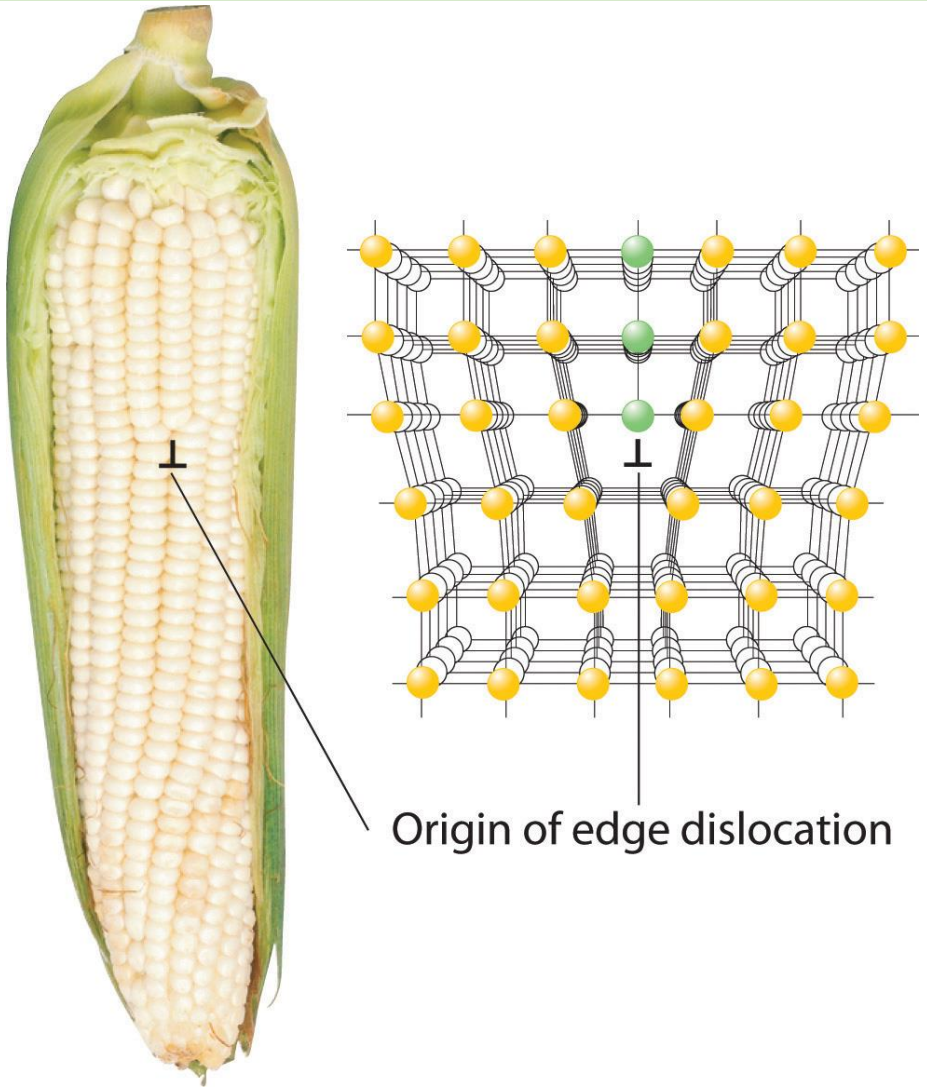
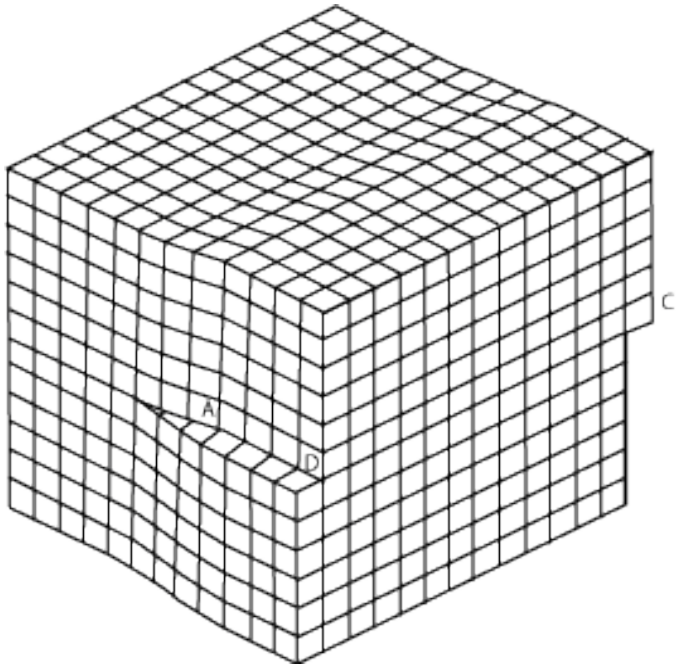
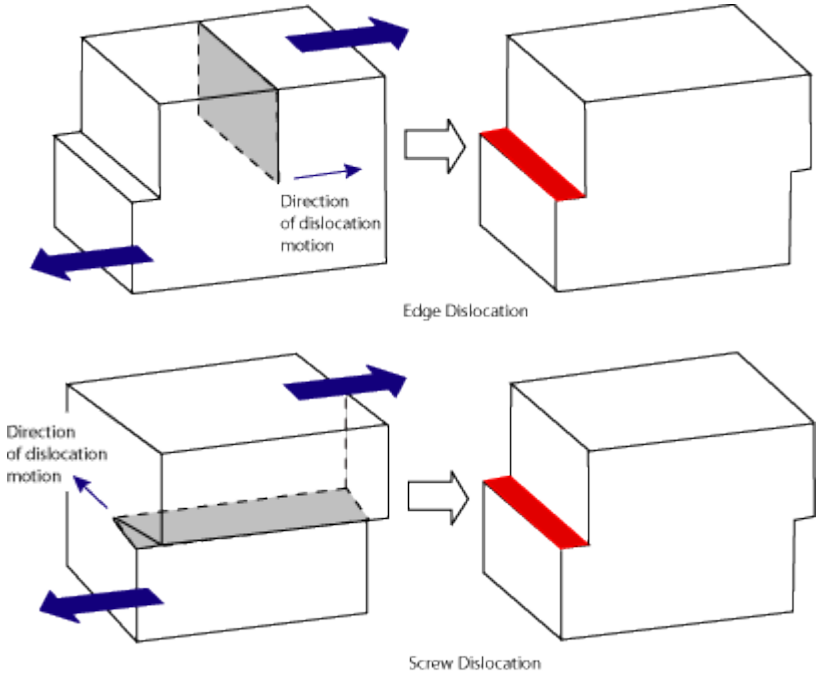
Crystal Defects



Crystal Defects





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